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What is the electronic configuration of potassium after it gets a +1 charge? Possible answers: Explanation: Potassium (K) is originally in the electronic configuration of [Ar]4s1. To obtain a charge +1 loses an electron, resulting in the configuration of [Ar]. Which of the following species is represented by the electronic configuration? Possible answers: Correct answer: Explanation: Due to the phenomenon of semiorbital stability in transition metals, electrons can easily move between 4s and 3d orbitals. The atom achieves greater stability from having only one atom in the 4s orbital, allowing a half-full 3d orbital, as opposed to a full 4s orbital and four electrons in the 3d subshell. For elements such as chromium and copper, which could have valence shell configurations of 4s23d4 and 4s23d9, respectively, an electron from the 4s orbital jumps down to the 3d orbital to take advantage of the stability added by the half-full orbital. The electronic configuration given is that of chrome. Note that you can also solve this issue by counting electrons to determine the atomic number. In this case, the electrons add up to 24, indicating the 24th element, chromium. What is the correct electronic structure for ? Possible answers: Correct answer: Explanation: If this question asked the electronic structure of magnesium (Mg) in its ground state, it would be the correct answer; however, the charge on means that the molecule emits two valence electrons to achieve a more stable orbital. Those electrons will be released from the outermost valence shell, which in this case is the shell; therefore, it is correct. Note that the ground state of magnesium will have twelve electrons (the same as its atomic number), while the ion will have ten. An oxygen atom and a fluorine atom have the same electronic configuration as neon. Which of the following on the atoms? Is it true? Possible answers: both atoms have the same charge because they both got the same amount electron electron The atoms are classified as halogen the fluorine has the mate of the charge of oxygen because it earns the quantity of electrons, the oxygen has the mate of the charge of fluorine because it earns the quantity of electrons correct response: the fluorine has the mate of the charge of Oxygen because it earns the quantity of electrons explanation: the question states that oxygen and fluorine atoms have the same electronic neon configuration. The electronic configuration for Neon is, with ten total electrons and eight Valenza electrons. To correspond to this configuration, both the oxygen atom and the fluorine atom must have eight electrons of value (two in the orbital and six in orbital). Recall that atoms in the oxygen group have six electrons of value, while halogens have seven electrons of value; Therefore, oxygen has six and the fluorine has seven Valenza electrons. To reach a total of eight eights of Valenza (Byte), the oxygen atom must have acquired two electrons, giving it a charge. Likewise, the fluorine must have acquired an electron, giving it a charge. The fluorine has the mate of the office of oxygen because it has gained the mate of the quantity of electrons to complete its octo. Which of the following statements is true about the gain and loss of electrons in cobalt? Possible answers: in the cobalt, the last orbital that earns electrons and the first orbital that loses electrons in cobalt, the last orbital that loses electrons in cobalt, the last orbital that earns electrons and the first orbital that loses electrons in cobalt is the first orbital that loses electrons in cobalt is the last orbital that loses correct response: the last orbital who earns electrons in cobalt is and the first orbital that He loses electrons in cobalt is explanation: cobalt is a transition metal, so it is found in block D periodic table. A cobalt atom on Earth has a configuration of electrons equal to . The last orbitals gaining or losing electrons shall be: orbital, since they are the orbital with the greatest energy and located further away from the nucleus. Remember that electrons are filled by low-energy to high-energy orbitals. An orbital has less energy than an orbital. This means that when you fill the electrons, the last orbital you fill is the orbital. When electrons are assigned to each orbital, two electrons are assigned to the orbital and then the remaining seven electrons are assigned to the five orbitals, so the orbitals are filled last when electrons are obtained. When an element loses electrons, the first orbital to lose electrons is the outermost one. This is because the adhesive force of the nucleus on the electron will be weaker in the outermost orbital (because it is farther away from the nucleus); therefore, it will be easy to move the electron away from the nucleus. In cobalt, the outermost orbital is the orbital (as it has the highest number of shells). This means that the electrons will be lost from the orbital before the orbital. Note that the orbital that gains electrons last is not the same orbital that loses electrons first. The acquisition of electrons depends on the energy of the orbital, and the loss of electrons depends on the location of the orbital. The most energetic orbital will gain the electron last and the outermost orbital will lose the electron first. A certain transition metal has an electronic configuration such that its orbital has only one electron. What would be a valid conclusion on this point? Possible answers: The identity of the element is chromium and contains one electron in each of its orbitals This is an invalid electronic configuration because the orbitals always contain two electrons in the transition metals The identity of the element is manganese and contains an empty orbital Correct answer: identity ilatibro ilatibro ious ied onucaic ni enorttele nu eneitnoc e omorc ²À eud I .inorttele eud osrep ah acirac anu ah iuc id omota nU .enoizageipS .atterroc atsopisR .etsopisr ilibissoF ?atterroc acimorttele enoizarugifnoc al ²À lauQ .oicsug otrauq len ilatot aznelav id inorttele euqnic ah ocinesra L .aznelav id enorttele olos nu ah emar II .elatibro aigrene atla ²Àip orol allen aznelav id inorttele eud litut onnah ociz e oisengam .oiclac .ocinesra .emarc .ocinez .emarc .ocinez .oisengam .oiclaC .inorttele id enoizarugifnoc alled jota ²Àip otaremun elatibrot onretnse oicsug len itaiggolla onnars aznelav id inorttele ilg .enoizageipS .ilbissop etsopisR ?aznelav id inorttele ²Àip ah otneleme elauQ .ilatibro ious ied onucaic ni enorttele nu e elatibro ous len enorttele nu Àrva onetsgmt li e ilatibro ious ied onucaic ni enorttele nu e elatibro ous len enorttele nu Àrva onedbihom II .) 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Àtem a itipmeir onos ilatibro ilg ittut odnauq elibats ²Àip ²À otneleme nu ehc amreffa oicsug ozzem a Àtilibats aL .oicsug ozzem a Àtilibats otamaihc onemonef nu id asuac a acifirev is ²ÀiC .ilatibro orol ied onu ni onatsops ol e elatibro orol lad inorttele ilged onu onodrep enoizisnart id illatem inucla .aivattuT .elatibro ous len inorttele eud onnah otilos id enoizisnart id illatem i ehc omaidrocir .elatibro ous len enorttele nu olos ah e enoizisnart id otneleme nu ²À otneleme'l ehc amreffa adnamod aL inorttele inorttele id enoizarugifnoc elamron anu ah muitnortS .alocelom allad iserp etnemlicaf ²Àip onos ehc .aznelav id inorttele ilg erpmes onos 38 electrons. In abbreviated notation, this would be: Lose two valence electrons removeÀ ⁵s electrons. This leaves the configuration as: This makes sense because it has a stable octet of valence. Strontium À is stable as an ion, which means it will have even an octet, allowing it to match the configuration of krypton. Which two elements have the same number of electrons in the 3d shell? Possible answers: Correct answer: Cr and Mn Explanation: The valence electrons in the subsoil d may be strange because a semi-filled orbital d ⁵ is more stable than one with three or four electrons. While strontium, titanium and vanadium have two electrons in their 4s orbital, chromium has one in the 4s shell and puts five in the 3d shell. This creates a semi-filled 4s and 3d shell, which are more stable than a full 4s shell and a partial 3d shell. Chromium and manganese then have the same number of 3d electrons: five. An atom with 1s22s22p6 electronic configuration could be one of the following, except . Answers: Explanation: This particular configuration denotes a particle with ten total electrons. The sodium atom, with eleven electrons, Àthe only listed one that could not have this configuration. Ionized sodium, however, symbolized by Na+, applies. (Be careful to distinguish neutral atoms and ions.) Gregory Certified Tutor of Pennsylvania State University-Main Campus, Bachelor of Science, Chemical Engineering, Carnegie Mellon University, Doctor of Ph... Andrew Certified Tutor University of South Florida-Main Campus, Bachelors, Biomedical Sciences. New York Medical College, Ph.D. Doctor of Medicine. Kevin Certified Tutor University of Nevada-Las Vegas, Bachelors, Mathematics. University Nevada-Las Vegas, graduate student, Radiochemistry. If you have found a problem with this question, please let us know. With help the community we can improve our educational resources. If you believe that content available through the website (such as In our terms of service) Viola one or more of your copyright, please notify us by providing a written notice ("NOTICE INFRINGEMENT" containing the information described following the designated agent listed below. 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